



## NAMES, FORMULAE, AND CHARGES OF SOME COMMON IONS

\* *Aqueous solutions are readily oxidized by air.*

\*\* *Not stable in aqueous solutions.*

<b>Positive Ions (Cations)</b>			
$\text{Al}^{3+}$	Aluminum	$\text{Pb}^{4+}$	Lead(IV), plumbic
$\text{NH}_4^+$	Ammonium	$\text{Li}^+$	Lithium
$\text{Ba}^{2+}$	Barium	$\text{Mg}^{2+}$	Magnesium
$\text{Ca}^{2+}$	Calcium	$\text{Mn}^{2+}$	Manganese(II), manganous
$\text{Cr}^{2+}$	Chromium(II), chromous	$\text{Mn}^{4+}$	Manganese(IV)
$\text{Cr}^{3+}$	Chromium(III), chromic	$\text{Hg}_2^{2+}$	Mercury(I)*, mercurous
$\text{Cu}^+$	Copper(I)*, cuprous	$\text{Hg}^{2+}$	Mercury(II), mercuric
$\text{Cu}^{2+}$	Copper(II), cupric	$\text{K}^+$	Potassium
$\text{H}^+$	Hydrogen	$\text{Ag}^+$	Silver
$\text{H}_3\text{O}^+$	Hydronium	$\text{Na}^+$	Sodium
$\text{Fe}^{2+}$	Iron(II)*, ferrous	$\text{Sn}^{2+}$	Tin(II)*, stannous
$\text{Fe}^{3+}$	Iron(III), ferric	$\text{Sn}^{4+}$	Tin(IV), stannic
$\text{Pb}^{2+}$	Lead(II), plumbous	$\text{Zn}^{2+}$	Zinc

<b>Negative Ions (Anions)</b>			
$\text{Br}^-$	Bromide	$\text{OH}^-$	Hydroxide
$\text{CO}_3^{2-}$	Carbonate	$\text{ClO}^-$	Hypochlorite
$\text{ClO}_3^-$	Chlorate	$\text{I}^-$	Iodide
$\text{Cl}^-$	Chloride	$\text{HPO}_4^{2-}$	Monohydrogen phosphate
$\text{ClO}_2^-$	Chlorite	$\text{NO}_3^-$	Nitrate
$\text{CrO}_4^{2-}$	Chromate	$\text{NO}_2^-$	Nitrite
$\text{CN}^-$	Cyanide	$\text{C}_2\text{O}_4^{2-}$	Oxalate
$\text{Cr}_2\text{O}_7^{2-}$	Dichromate	$\text{O}^{2-}$	Oxide**
$\text{H}_2\text{PO}_4^-$	Dihydrogen phosphate	$\text{ClO}_4^-$	Perchlorate
$\text{CH}_3\text{COO}^-$	Ethanoate, acetate	$\text{MnO}_4^-$	Permanganate
$\text{F}^-$	Fluoride	$\text{PO}_4^{3-}$	Phosphate
$\text{HCO}_3^-$	Hydrogen carbonate, bicarbonate	$\text{SO}_4^{2-}$	Sulphate
$\text{HC}_2\text{O}_4^-$	Hydrogen oxalate, binoxalate	$\text{S}^{2-}$	Sulphide
$\text{HSO}_4^-$	Hydrogen sulphate, bisulphate	$\text{SO}_3^{2-}$	Sulphite
$\text{HS}^-$	Hydrogen sulphide, bisulphide	$\text{SCN}^-$	Thiocyanate
$\text{HSO}_3^-$	Hydrogen sulphite, bisulphite		